Medical Faculties ((Pharmacy, Medicine, Dentistry, Health sciences))

Admission Exam Sample

General Instructions

1- The First Page of the booklet is the answer sheet. Fold this page along the perforations, slowly and carefully tear off the answer sheet.

2- Write your name and your seat number then fill the seat number in the proper place on the answer sheet.

3- Be sure to fill only one answer with a pencil for each question.

4- Use of mobiles is prohibited
1. Number of functional groups in the following compound is

![Chemical structure](image)

a. 1  
b. 2  
c. 3  
d. 4

2. Which two elements react together to form an ionic compound?

<table>
<thead>
<tr>
<th>element</th>
<th>Electronic structure</th>
</tr>
</thead>
<tbody>
<tr>
<td>W</td>
<td>2,4</td>
</tr>
<tr>
<td>X</td>
<td>2,8</td>
</tr>
<tr>
<td>Y</td>
<td>2,8,1</td>
</tr>
<tr>
<td>Z</td>
<td>2,8,7</td>
</tr>
</tbody>
</table>

a. W and X  
b. X and Y  
c. Y and Z  
d. Z and W
3. Which diagram shows the process of diffusion?

A

B

C

D

key

○ different atoms

4. The diagram shows that two gases are formed when concentrated hydrochloric acid is electrolysed using inert electrodes.

Which row correctly describes the colours of the gases at the electrodes?

<table>
<thead>
<tr>
<th>Anode(+ve)</th>
<th>Cathode(-ve)</th>
</tr>
</thead>
<tbody>
<tr>
<td>a. colorless</td>
<td>colorless</td>
</tr>
<tr>
<td>b. colorless</td>
<td>Yellow-green</td>
</tr>
<tr>
<td>c. Yellow-green</td>
<td>colorless</td>
</tr>
<tr>
<td>d. Yellow-green</td>
<td>Yellow-green</td>
</tr>
</tbody>
</table>
5. Which method is most suitable to obtain zinc carbonate from a suspension of zinc carbonate in water?

a. crystallisation  
b. distillation  
c. evaporation  
d. filtration

6. The diagram shows the reaction between zinc oxide and dilute hydrochloric acid.

Which terms describe the reaction?

<table>
<thead>
<tr>
<th></th>
<th>endothermic</th>
<th>neutralization</th>
</tr>
</thead>
<tbody>
<tr>
<td>a.</td>
<td>✓</td>
<td>✓</td>
</tr>
<tr>
<td>b.</td>
<td>✓</td>
<td>X</td>
</tr>
<tr>
<td>c.</td>
<td>X</td>
<td>✓</td>
</tr>
<tr>
<td>d.</td>
<td>X</td>
<td>X</td>
</tr>
</tbody>
</table>

7. The element vanadium, V, forms several oxides. In which change is oxidation taking place?

a. $\text{VO}_2 \rightarrow \text{V}_2\text{O}_3$  
b. $\text{V}_2\text{O}_3 \rightarrow \text{VO}_2$  
c. $\text{V}_2\text{O}_3 \rightarrow \text{VO}$  
d. $\text{V}_2\text{O}_3 \rightarrow \text{V}_2\text{O}_5$

8. A gas is escaping from a pipe in a chemical plant. A chemist tests this gas and finds that it is alkaline. What is this gas?

a. ammonia  
b. chlorine  
c. hydrogen  
d. sulfur dioxide
9. The graph shows how the pH changes as an acid is added to an alkali.

\[
\text{acid + alkali} \rightarrow \text{salt + water}
\]

Which letter represents the area of the graph where both acid and salt are present?

![Graph showing pH changes](attachment:image.png)

10. Select the right arrangement to describe the cleansing action of soaps:

![Diagram showing cleansing action of soaps](attachment:image.png)

a. a, b, c, d  
b. a, b, d, c  
c. a, c, b, d  
d. a, d, b, c

11. Name the following reaction:

\[
\begin{align*}
\text{RCOO}_2R' + 3 \text{NaOH} & \overset{H_2O}{\rightarrow} \text{HOCH}_2\text{OH} + 3 R'\text{COO}^\ominus \text{Na} \\
\end{align*}
\]

a. Denaturation  
b. Hydroxylation  
c. Esterification  
d. Saponification
12. One of the following structures is an antibiotic:

\[ \text{a. Compound a} \]
\[ \text{b. Compound b} \]
\[ \text{c. Compound c} \]
\[ \text{d. Compound d} \]

13. Phosphoric acid is a .......... compound:

\[ \text{a. Monoprotic} \]
\[ \text{b. Diprotic} \]
\[ \text{c. Triprotic} \]
\[ \text{d. Tetraprotic} \]

14. Which of the following structures is cyclohexane?

\[ \text{a. Structure a} \]
\[ \text{b. Structure b} \]
\[ \text{c. Structure c} \]
\[ \text{d. Structure d} \]

15. The loss of a secondary, tertiary, or quaternary protein structure due to the disruption of monovalent interactions and/or disulfide bonds that leaves the primary structure intact is known as:

\[ \text{a. Esterification} \]
\[ \text{b. Denaturation} \]
\[ \text{c. Saponification} \]
\[ \text{d. Dehydration} \]
Study the following structure then choose the correct answer for questions 16-20:

![Chemical structure](image)

16. Number of sp carbons:
   a. 2
   b. 3
   c. 6
   d. 4

17. Number of sp\(^3\) carbons
   a. 2
   b. 3
   c. 6
   d. 4

18. Number of sp\(^3\) carbons
   a. 2
   b. 3
   c. 6
   d. 4

19. Number of \(\sigma\)-bonds
   a. 8
   b. 5
   c. 16
   d. 4
20. Number of \( \pi \)-bonds
   
a. 2  
b. 3  
c. 6  
d. 4

21. Which test could be used to find which end of a magnet is the north pole?  
   a. putting it near a compass needle  
   b. putting it near a ferrous metal  
   c. putting it near a non-ferrous metal  
   d. putting it near a steel spoon

22. In two separate experiments, a magnet is brought near to an unmagnetised iron bar. This causes the bar to become magnetised

   \[\text{experiment 1} \quad \begin{array}{c} \text{N} \quad \text{magnet} \quad S \\ \rightarrow \end{array} \quad \begin{array}{c} X \quad \text{iron bar} \end{array}\]

   \[\text{experiment 2} \quad \begin{array}{c} S \quad \text{magnet} \quad N \\ \rightarrow \end{array} \quad \begin{array}{c} \text{iron bar} \quad Y \end{array}\]

Which magnetic poles are induced at \(X\) and at \(Y\)?

<table>
<thead>
<tr>
<th></th>
<th>Pole induced at (X)</th>
<th>Pole induced at (Y)</th>
</tr>
</thead>
<tbody>
<tr>
<td>a</td>
<td>N</td>
<td>N</td>
</tr>
<tr>
<td>b</td>
<td>N</td>
<td>S</td>
</tr>
<tr>
<td>c</td>
<td>S</td>
<td>N</td>
</tr>
<tr>
<td>d</td>
<td>S</td>
<td>S</td>
</tr>
</tbody>
</table>
23. A student sets up the circuit shown. The switch is open (off).

Which lamps are on and which lamps are off?

<table>
<thead>
<tr>
<th></th>
<th>Lamp X</th>
<th>Lamp Y</th>
<th>Lamp Z</th>
</tr>
</thead>
<tbody>
<tr>
<td>a</td>
<td>off</td>
<td>off</td>
<td>off</td>
</tr>
<tr>
<td>b</td>
<td>on</td>
<td>off</td>
<td>off</td>
</tr>
<tr>
<td>c</td>
<td>on</td>
<td>off</td>
<td>on</td>
</tr>
<tr>
<td>d</td>
<td>on</td>
<td>on</td>
<td>on</td>
</tr>
</tbody>
</table>

24. A powder contains 400 mg of a radioactive material that emits α-particles. The half-life of the material is 5 days. What mass of that material remains after 10 days?

a. 0 mg.
b. 40 mg
c. 100 mg
d. 200 mg

25. An ammeter and an 18 Ω resistor are connected in series with a battery. The reading on the ammeter is 0.50 A. The resistance of the battery and the ammeter can be ignored.

What is the electromotive force (e.m.f.) of the battery?

a. 9.0 N
b. 9.0 V
c. 36 V
d. 36 N
26. A polythene rod repels an inflated balloon hanging from a nylon thread. What charges must the rod and the balloon carry?

a - The rod and the balloon carry opposite charges
b - The rod and the balloon carry like charges
c - The rod is charged but the balloon is not
d - The balloon is charged but the rod is not

27. Which line in the table describes the properties of solids and of liquids at fixed temperature?

<table>
<thead>
<tr>
<th>Solids</th>
<th>Liquids</th>
</tr>
</thead>
<tbody>
<tr>
<td>a definite volume and definite shape</td>
<td>no definite volume but definite shape</td>
</tr>
<tr>
<td>b no definite volume but definite shape</td>
<td>definite volume and definite shape</td>
</tr>
<tr>
<td>c definite volume and definite shape</td>
<td>definite volume but no definite shape</td>
</tr>
<tr>
<td>d no definite volume but definite shape</td>
<td>no definite volume and no definite shape</td>
</tr>
</tbody>
</table>

28. The diagram shows a thermometer calibrated in degrees Celsius.

![Thermometer Diagram]

What are the values of the lower fixed point and of the upper fixed point on the Celsius scale?

<table>
<thead>
<tr>
<th></th>
<th>Lower fixed point/°C</th>
<th>Upper fixed point/°C</th>
</tr>
</thead>
<tbody>
<tr>
<td>a</td>
<td>-10</td>
<td>110</td>
</tr>
<tr>
<td>B</td>
<td>0</td>
<td>20</td>
</tr>
<tr>
<td>C</td>
<td>0</td>
<td>100</td>
</tr>
<tr>
<td>d</td>
<td>20</td>
<td>100</td>
</tr>
</tbody>
</table>
29. Sound travels by wave motion. Which property of waves causes echoes?
   a. diffraction
   b. dispersion
   c. reflection
   d. refraction

30. A thin converging lens is used to produce, on a screen, a focused image of a candle.

   Various focused images are produced on the screen by moving the lens and the screen backwards and forwards. Which statement is always correct?
   a. The image is at the principal focus (focal point) of the lens
   b. The image is bigger than the object
   c. The image is closer to the lens than the object is
   d. The image is inverted

31. The gonadotropin-releasing hormone is secreted by the
   a. Hypothalamus
   b. anterior pituitary
   c. posterior pituitary
   d. ovary

32. The contraceptive method depending on negative feedback effect
   a. intrauterine device
   b. oral contraceptive pill
   c. diaphragm
   d. condom
33. Women with blocked oviduct may get help by using
   a. Mifepristone
   b. Progesterone
   c. artificial insemination
   d. in vitro fertilization

34. Lipids are compounds formed of the following atoms
   a. C, H, O, N
   b. C, H, O
   c. C, H, N
   d. C, H, P

35. Starch is an example of
   a. Carbohydrates
   b. Proteins
   c. Lipids
   d. fatty substances

36. A daily energy expenditure value of an adult in Kj
   a. 70
   b. 700
   c. 7000
   d. 70000

37. A Mutation that can lead to premature chain termination
   a. Frame shift mutation
   b. Silent mutation
   c. Mis-sense mutation
   d. Non-sense mutation

38. In a dihybrid cross with linked genes, crossing-over occurs during meiosis, allowing the formation of gametes which are:
   a. Only parental
   b. Only recombinant
   c. Parental and recombinant
   d. Identical
39. If one the parents died from lung cancer, this does not put the person at a greater risk than others whose parents did not develop cancer. Explain how cancer is a disease caused by mutations and yet not heritable.
   a. Mutations caused by environmental agents such as tobacco smoke cannot be inherited.
   b. Mutations that cause cancer are special and cannot be passed on regardless of what type of cell they occur in.
   c. Most cancers arise from mutations in germ cell line cells.
   d. Most cancers arise from mutations in somatic cells.

40. If you are a pharmacist at the quality control unit and have been asked to identify a white, unlabeled powder. As a first step, you are interested in learning is polar or non-polar. What you do and how should you interpret your findings?
   a. Try to dissolve the powder in water; if it fails to dissolve, it is a polar compound.
   b. Try to dissolve the powder in water; if it fails to dissolve, it is a non-polar compound.
   c. Try to ignite the powder; if it burns, it is a polar compound.
   d. Try to dissolve the powder in oil; if it dissolves, it is a polar compound.

41. The addition of anti-SRBC rabbit antibodies to sheep red blood cells (SRBC), incubated for at 37°C and observed 1-2 hours later lead to:
   a. Sedimentation of the intact SRBC
   b. Sedimentation of the agglutinated SRBC
   c. No sedimentation of SRBC
   d. Hemolysis, destroyed SRBC

42. Virus is the smallest infective agent that belong to which type of cells
   a. Prokaryotes
   b. Eukaryotes
   c. Non prokaryotes- non eukaryotes
   d. Both prokaryotes and eukaryotes.
43. The next graph shows a serious of electrophoresis of the hemoglobin of seven individuals. The difference in migration are due to the differences in the amino acid sequences of the two polypeptide chains of β-globin, which form the hemoglobin molecule along with the two α-globin chains.

![Graph showing electrophoresis of hemoglobin]

How many possible alleles of the β-globin gene are shown by gel electrophoresis?

a. 2  
b. 3  
c. 4  
d. 5

44. The reintroduction of the same antigen into the body help create:

a. Primary response  
b. Latent response  
c. Secondary response  
d. Activates T cells only

45. T lymphocytes recognize antigens specifically through membrane bound T cell receptor. T helper lymphocytes belong to this group of cells that act to:

a. Kill the infected cells  
b. Express surface antibodies  
c. Secrete circulating antibodies  
d. Regulate the activity of other cells of immunity

46. The polymer of (C₆H₁₀O₅) molecules is called

a. Neoglucose  
b. Glucose  
c. Fructose  
d. Glycogen
47. A human can survive only a few hours after
   a. Fasting
   b. liver ablation
   c. diabetes mellitus
   d. glycogenolysis

48. A healthy individual after 1 hour of provoked hyperglycemia will have a blood
   glucose concentration of:
   a. 1.5 g/l
   b. 5.1 g/l
   c. 0.6 g/l
   d. 0.8 g/l

49. The female sexual hormones are secreted by the
   a. Uterus
   b. Endometrium
   c. Ovary
   d. Vagina

50. Hormone(s) secreted during the luteal phase
   a. Androgen
   b. Estrogen
   c. Progesterone
   d. estrogen and progesterone

51. Two women play five games of checkers. Each woman wins the same number of
   games. There are no ties. Explain this.
   a. Because they aren't wearing ties!
   b. The women are not playing each other.
   c. These two women are playing men.
   d. Women don't wear ties.

52. There are 3 apples and you take away 2. How many do you have?
   a. One
   b. Two
   c. Three
   d. Four

53. If you have only one match and you walked into a room where there was an oil
   burner, a kerosene lamp, and a wood burning stove, what would you light first?
   a. the kerosene lamp because it is easier
   b. the oil burner
   c. the wood burning stove.
   d. the thing in my hand
54. A forester has 17 trees, and all but 9 die. How many are left?
   a. 8 trees
   b. 9 trees
   c. the first 8
   d. I don't know

55. How many 2 cent stamps are in a dozen?
   a. Six
   b. Twelve
   c. Eighteen
   d. Twenty-four

56. Name the greatest of all inventors.
   a. Accident
   b. Gertrude Stein
   c. Einsten
   d. Alexander Graham Bell

57. What is the missing figure? 4, 5, 8, 17, 44.....
   a. 80
   b. 125
   c. 112
   d. 60

58. A young monk goes fishing with two other much older and experienced monks. While out in the boat one of the older monks realizes he has forgotten his favorite hook, gets out of the boat, walks across the water, and returns shortly with his hook. The young monk is stunned at this sight. A short while later the other monk needs to recover a fishing cork he has dropped into the lake. He gets out of the boat, walks across the water about 15 feet, picks up the cork, and returns to the boat. The young monk is wide eyed at this miracle, knowing full well that what he thought impossible can be done. Later the young monk makes up an excuse to go ashore so he steps out of the boat, sinks over his head in the very cold water, and comes bobbing and sputtering to the surface. The young monk asks the two old monks, "How did you do walk on water?"
   a. The first monk replied, "you are not holy enough".
   b. The second monk replied, "it is easy if you know where the rocks are."
   c. Neither monk had an explanation for their skill. You can walk on water or you can't.
   d. That is the way it is.
59. Sue: Commercial flights currently contribute more carbon dioxide to the atmosphere in one year than does the whole of Africa. If we want to reduce global warming we need to restrict the number of flights we take.
Dave: Did you know that by taking one inter-continental flight you cause more pollution than you would in twelve months of car travel?
Dave’s response to Sue’s comment serves to:
   a. reinforce Sue’s contention that flights are a major contributor to increased carbon dioxide levels
   b. add more weight to her contention that we should reduce the number of flights we take
   c. mitigate the force of her argument by suggesting that there is an alternative approach
   d. suggest an alternative that will reduce the effect of pollution

60. Josh has twenty years of typing experience behind him; therefore, if you are looking for an efficient typist to enter your data into the new system, you need look no further. The speaker assumes that:
   a. Twenty years of practice ensures typing efficiency
   b. The type of typing required for the new system is identical to what Josh has been doing
   c. Josh’s job profile is the best that the new employer is going to get
   d. Josh is an outstandingly fast and accurate typist

61. What is the missing figure? 13, 57, 911, 1315, 1719....
   a. 2123
   b. 1879
   c. 3002
   d. 5004

62. It has been suggested that long-term prisoners, on release from jail, be given a reasonable state pension to reduce the likelihood of their resorting to crime. Most people instinctively reject the suggestion as they feel it would be like rewarding criminal activity.
The supporters of the prisoners’ pension scheme have criticized those who reject this possibility, by claiming that for the critics
Which of the following is the most logical completion of the sentence above?
   a. emotion is more important than justice
   b. punishment for criminals is more important than crime prevention
   c. crime prevention is not an important issue
   d. money has too high a value
63. Pick the piece that's missing from the diagram below

![Diagram with options a, b, c, d]

64. Find the picture that follows logically from the diagrams to the right.

![Diagram with options A, B, C, D]

65. How many four sided shapes does this diagram have?

a. 5-10
b. 11-15
c. 16-20
d. 26-30
66. You are trying to decide what kind of car to buy. You make a chart to compare a two-seater sports car, a two-door sedan, and a minivan in three categories. Which would NOT be a suitable category?
   a. price
   b. gas mileage
   c. tire pressure
   d. storage capacity

67. Which is NOT a likely cause of this situation?
   a. "I can't turn on the lamp in the family room!"
   b. The lamp isn't plugged into an electrical outlet.
   c. We just bought a new couch in a color that matches the lamp.
   d. The light bulb in the lamp has burned out

68. How many birthdays does the average man have?
   a. 78
   b. 65.5
   c. one per year
   d. None

69. Some months have 31 days, how many have 28?
   a. One month and that is February
   b. February, every 4 years.
   c. All months have (at least) 28 days
   d. None of the above

70. A woman gives a beggar 50 cents, the woman is the beggar's sister, but the beggar is not the woman's brother. How come?
   a. The beggar is the woman's sister
   b. Not possible.
   c. The beggar is an in-law.
   d. The woman is the beggar.
Pharmacy Oriented Questions

71. Food supplements include the following, EXCEPT:
   a. Aspirin
   b. Ascorbic acid
   c. Vitamin A
   d. Omega-3-fatty acids

72. Osteoporosis can be treated with a combination of calcium and:
   a. Vitamin A
   b. Vitamin B
   c. Vitamin C
   d. Vitamin D

73. Anemia can be treated by the use of:
   a. Iron
   b. Potassium
   c. Copper
   d. Calcium

74. Among the internally used medications are:
   a. Ointments
   b. Capsules
   c. Creams
   d. Patches

75. Which of the following vitamins is used for the treatment of night blindness?
   a. Vitamin A
   b. Vitamin B
   c. Vitamin D
   d. Vitamin C

76. If a tablet that can be easily divided into two halves, contains 500 mg of a drug and the patient should take 750 mg of that drug daily, then how many tablets would you recommend for a one-week use?
   a. 11.5
   b. 12
   c. 9.5
   d. 10.5

77. The recommended medication for a person suffering from an insect bite is more likely to be an:
   a. Anti-allergic medicine
   b. Anti-pyretic medicine
   c. Analgesic medicine
   d. Anti-rheumatic medicine
78. The doctor’s prescription should contain all of the following information EXCEPT:
   a. The patient’s name
   b. The pharmacy’s name
   c. The drug’s name
   d. The doctor’s name

79. Anticancer medications (chemotherapy) can cause:
   a. Constipation
   b. Fever
   c. Hair loss
   d. Hypertension

80. If you are home and don’t have access to medications and you were suffering from a heartburn (hyperacidity), which of the following household substances would you use to reduce your heartburn?
   a. Soda beverages
   b. Milk
   c. Lemon juice
   d. Vinegar

81. Sunscreen creams are used to protect the skin against:
   a. Infrared radiations
   b. Microwaves
   c. Ultra violet radiations
   d. Ultrasound waves

82. Medications should be stored at home away from:
   a. Humidity
   b. Noise
   c. Electric source
   d. Electronic devices

83. Which of the following tasks is NOT a pharmacist’s responsibility?
   a. Dispensing medicines
   b. Counselling on medicines
   c. Diagnosis of diseases
   d. Communicating medical information

84. Tonsillitis can be treated by an.................medication:
   a. Anti-viral
   b. Antibacterial
   c. Anti/protozoal
   d. Anti-fungal

85. Medications should be kept away from:
   a. Newborns
   b. Children
   c. Teenagers
   d. Old people
86. In case of fire burns, the first aid approach should be to:
   a. Cover the burn with oil
   b. Wash the burn with warm water
   c. Wash the burn with cold water
   d. Cover the burn with bandage

87. A person should never take an old stored medication before checking the:
   a. Name of Pharmacy where the medication was purchased
   b. Expiry date of the medication
   c. Name of doctor who prescribed the medication
   d. Number of doses of the medication per day

88. Vaccines can be used to protect against:
   a. Influenza
   b. Diabetes
   c. AIDS
   d. Allergy

89. It is recommended for children to take medications in the form of:
   a. Injections
   b. Tablets
   c. Capsules
   d. Syrups

90. Which of the following drugs can be given to patients without a physician’s prescription?
   a. Antibiotics
   b. Analgesics
   c. Anticancer drugs
   d. Antidepressants
Medicine & Health Sciences Oriented Questions

91. Mosquitos may transmit
   a. Malaria.
   b. Toxoplasmosis.
   c. Anorexia nervosa.
   d. Scurvy.

92. Sexually transmitted diseases include all the following, EXCEPT
   a. Gonorrhea.
   b. AIDS.
   c. Rheumatic fever.
   d. Syphilis.

93. What is the largest gland of the human body?
   a. The pancreas.
   b. The liver.
   c. The parotid salivary gland.
   d. The skin.

94. What is anemia?
   b. Rapid death of all blood cells.
   c. Decreased red blood cells.
   d. Cancer of the white blood cells.

95. An American-Lebanese cardiac surgeon who developed a machine that made open heart surgery possible
   a. Michael Dabaghi.
   b. Magdi Yacoub.
   c. Christiaan Barnard.
   d. Joseph Woo.

96. Influenza is a ________ disease
   a. bacterial.
   b. viral.
   c. protozoal.
   d. autoimmune.
97. Who composed “Kitab Al-Qanun Fi Al-tibb”?
   b. Avicenna (Ibn Sina).
   c. Abu Al-Qasim Al-Zahrawi.
   d. Abu- Bakr Al-Razi.

98. Today, medical students swear an oath upon graduating which is derived from the original one written by
   a. Imhotep.
   b. Aristo.
   c. Ikhnaton.
   d. Hippocrates.

99. Hyperglycemia means
   a. Increased pressure of the fluid filling the eye.
   b. Increased blood sugar.
   c. Increased arterial blood pressure.
   d. Increase fluid in the tissues of the body.

100. Nobel’s prizes were established in 1895 by the will of the Swedish chemist Alfred Nobel who invented
    a. optic fibers.
    b. polyethylene.
    c. penicillin.
    d. dynamite.

101. The first woman to win a Nobel Prize was
    a. Marie Curie.
    b. Grace Hopper.
    c. Kate Gleason.
    d. Fiona Wood.

102. The average volume of the blood in a normal adult is about
    a. 5 Liters.
    b. 7 Liters.
    c. 9 Liters.
    d. 11 Liters.
103. Neil Armstrong
   a. invented the light microscope.
   b. was the first pilot to cross the Atlantic ocean.
   c. is an American actor.
   d. was the first man to walk on the moon.

104. A skydiver, who was the first person to break the sound barrier without vehicular power
   b. Felix Baumgartner.
   c. Jay Stokes.
   d. Glenn Singleman.

105. The mineral present in the human body in largest amount is
   a. iron.
   b. potassium.
   c. calcium.
   d. copper.

106. Patients with the disease Diabetes Mellitus may need injection treatment with
   a. adrenalin.
   b. insulin.
   c. thyroxin.
   d. gastrin.

107. Which of the following vitamins is most important for healthy bone?
   a. Vitamin K.
   b. Vitamin B.
   c. Vitamin D.
   d. Vitamin A.

108. The headquarters of the United Nations (U.N) is located in
   a. Rome.
   b. New York.
   c. Geneva.
   d. London.
109. The red color of the blood is due to the presence of
   a. hemoglobin.
   b. bilirubin.
   c. cholesterol.
   d. vitamin B.

110. The main gas in the atmospheric air is
   a. oxygen.
   b. carbon dioxide.
   c. nitrogen.
   d. helium.